

# **Experiment Nein! Acid-Base Equilibria**

**CH 204**

**Fall 2008**

**Dr. Brian Anderson**

# Last Week

## Heat in chemical reactions:

heat is a measurable quantity

produced and consumed in stoichiometric amounts

## Heat Capacity (J/K):

how much heat is required to raise the temperature of something by one degree Celsius (or 1 Kelvin)

## Specific Heat Capacities (J/gK):

Lead 0.128

Iron 0.449

Water 4.184

## Hess's Law

# **This week**

**Weak acid titration.**

**Determine  $K_a$  of acetic acid by a couple different methods.**

**Witness the awesome power of a buffer solution to resist changes in pH.**

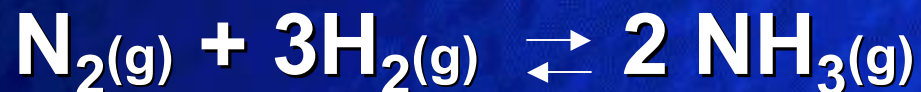
# Non-Equilibrium Reaction

Reaction goes to completion.



# Equilibrium Reaction

Products react with each other to re-form the reactants.



# Equilibrium Expression

For any equilibrium reaction,



products  
reactants

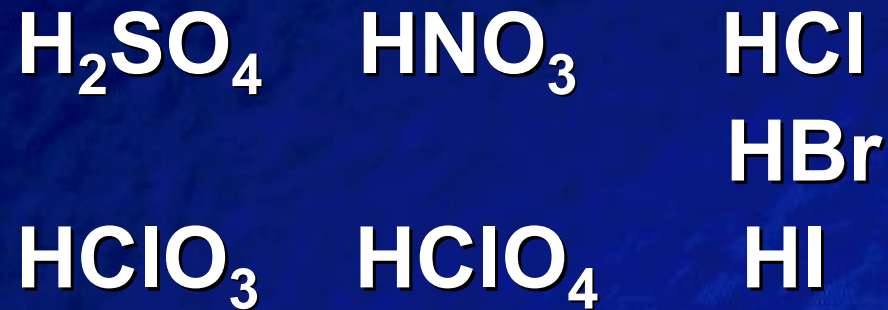
$$K_{\text{eq}} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

For a weak acid dissociation,



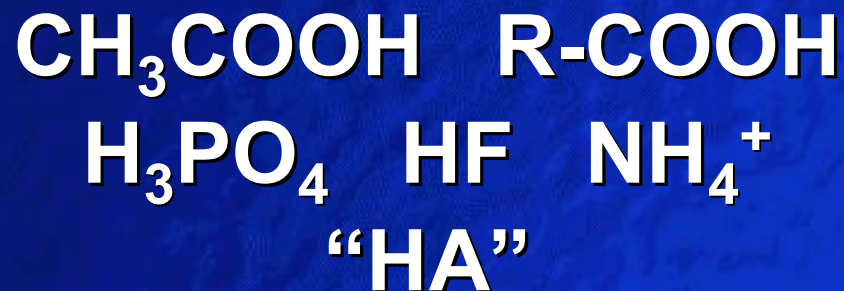
$$K_a = \frac{[H^+][A^-]}{[HA]}$$

# Strong Acids



# Weak Acids

*All the rest!*



# Calculating $[H^+]$ and pH

If you are given the  $[H^+]$

$$pH = -\log[H^+]$$

If you are given the pH

$$[H^+] = 10^{-pH}$$



# Four-Part Lab

## 1) Calibrate pH meter

Make sure you're in CALIBRATION mode.

Calibrate the pH meter in the order in the lab manual: **pH 7** first, then **pH 4**, then **pH 10**.

Press ENTER or CON to confirm calibration.

# Last Two Titrations of Your Life

## 2) Titrate 25 ml of 0.1 M acetic acid using 0.1 M NaOH

DO NOT add water! No indicator this time. Titrate in a beaker, not a flask, because you need room for the pH electrode.

Record pH after the addition of every 1.0 ml of NaOH at first, and as the pH begins to change more quickly, record smaller volume increments, down to 0.2 or 0.1 ml. Try to catch points on the vertical portion of the graph.

Switch roles with your lab partner and repeat the titration a second time.

Graph pH (y-axis) versus ml added (x-axis) in Excel.

# At the Equivalence Point

All of the HA has been reacted away.

If the solution was initially 0.1M acetic acid,  
it is now 0.05 M sodium acetate

At the half-equivalence point, half of the HA  
has been reacted away, and the [HA] and  
[A<sup>-</sup>] concentrations are equal.

# Half-Equivalence Point

At the half-equivalence point,  $[HA] = [A^-]$ .

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

$$K_a = [H^+] \times \frac{[A^-]}{[HA]}$$

$$-\log K_a = -\log [H^+]$$

$$\text{pH} = \text{p}K_a$$

# A Short-Cut to $K_a$

- 3) Measure the pH of 1.2 M acetic acid and two buffer solutions

Use measured pH and known acetic acid and acetate concentrations to calculate  $K_a$

$$K_a = \frac{[10^{-\text{pH}}][A^-]}{[HA]}$$

# Dissociation of acetic acid

	$\text{CH}_3\text{COOH}$	$\text{CH}_3\text{COO}^-$	$\text{H}^+$
initial	1.2	0	0
equilibrium	$1.2 - x$	$x$	$x$

$$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} = \frac{[x][x]}{[1.2 - x]}$$

$$x = 10^{-\text{pH}}$$

Assume  $x \ll 1.2 \text{ M}$

# Dissociation of acetic acid

If  $x \ll 1.2$  M, then

$$K_a = \frac{[x][x]}{[1.2 - x]} = \frac{[x^2]}{[1.2]}$$

$$K_a = \frac{(10^{-\text{pH}})^2}{[1.2]}$$

# Buffer Solutions



When you mix HA and A<sup>-</sup>, very little dissociation takes place (Le Chatlier's Principle).

So the number of moles of HA and of A<sup>-</sup> remain essentially constant.

Mixing solutions of HA and A<sup>-</sup> will dilute both of them.

Calculate the new [HA] and [A<sup>-</sup>] after the dilution, then use that value in the equilibrium expression.



# And Finally...

- 4) Add strong acid & base to buffers and to water and compare the changes in pH.

# pH meters need love, too

**Glass bulb is very thin**

**Remove carefully from storage bottle – turn the bottle, not the cap**

**Rinse well between samples, dab dry with KimWipes**

**Keep the bulb wet between readings**

**Swish samples to get better readings**



# Next week

**Final Class Meeting**

**Course/Instructor Surveys**

**TA Evaluations**

**Kinetics lab**

**Lab check-out**

**If you have missed more than one lab,  
e-mail me to discuss make-up week.**

# Next week's quiz

Given  $[H^+]$  calculate pH

Given pH, calculate  $[H^+]$

Know how to recognize a buffer solution

Know how to make up a buffer solution

Given three variables in an equilibrium expression, calculate the fourth. (Also, concentration is moles per liter.)

Dilution problems.

# **This week's quiz**

**Calorimetry and Hess's Law,  
just like I promised last week.**

**Only one more quiz after this one.**

**No lecture after Experiment 10.**

**No quiz after Experiment 10.**

**Make-up labs November 24 & 25**